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Balancing Equations and Simple Stoichiometry-KEY
Balance the following equations: 1) $1 \text{ N}_2 + 3 \text{ F}_2 \rightarrow 2 \text{ NF}_3$ 2) $2 \text{ C}_6\text{H}_{10} + 17 \text{ O}_2 \rightarrow 12 \text{ CO}_2 + 10 \text{ H}_2\text{O}$ 3) $1 \text{ HBr} + 1 \text{ KHCO}_3 \rightarrow \dots$ The smaller of these two answers is correct, and the reagent that leads to this answer is the limiting reagent. Both

Balancing Equations

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(Coefficients equal to one

(1) do not need to be shown in your answers).

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composed of matter; you, me

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and everything around us is made up of matter. Matter is a physical substance that takes up space and has mass.

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Balancing chemical equations means that you write the chemical equation correctly so that there is the same amount of mass on each side of the arrow. In this section, we're going to explain how to balance a chemical equation by using a real life example, the chemical equation that occurs when iron rusts:

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